

# Chemistry – Course Overview

<b>1<sup>st</sup> QUARTER:</b>	<b>2<sup>nd</sup> QUARTER:</b>
<p data-bbox="604 699 785 724"><b>BENCHMARK I</b></p> <p data-bbox="594 732 795 764"><b>3<sup>rd</sup> QUARTER:</b></p> <p data-bbox="386 802 638 826"><b>X. Chemical Quantities</b></p> <ul data-bbox="436 829 999 911" style="list-style-type: none"><li>A. The Mole – measuring matter</li><li>B. Mass &amp; Mole – Volume Relationship</li><li>C. Percent Composition &amp; Chemical Form - Mathematical</li></ul> <p data-bbox="289 914 411 938">Calculations</p> <p data-bbox="386 941 638 966"><b>XI. Chemical Reactions</b></p> <ul data-bbox="436 969 772 1050" style="list-style-type: none"><li>A. Describing Chemical Reactions</li><li>B. Types of Chemical Reactions</li><li>C. Reactions in aqueous solutions</li></ul> <p data-bbox="386 1053 588 1078"><b>XII. Stoichiometry</b></p> <ul data-bbox="436 1081 873 1187" style="list-style-type: none"><li>A. Mole Problems only simple stoichiometry<ul data-bbox="487 1105 852 1187" style="list-style-type: none"><li>1. Arithmetic of Equations</li><li>2. Chemical Calculations</li><li>3. Limiting reagent and percent yield</li></ul></li></ul> <p data-bbox="594 1487 795 1511"><b>BENCHMARK III</b></p>	<p data-bbox="1419 699 1600 724"><b>BENCHMARK II</b></p> <p data-bbox="1419 732 1621 764"><b>4<sup>th</sup> QUARTER:</b></p> <p data-bbox="1220 802 1472 826"><b>XIII. States of Matter</b></p> <ul data-bbox="1270 829 1740 911" style="list-style-type: none"><li>A. Nature of Gases &amp; Nature of Solids</li><li>B. Changes of State</li><li>C. Intermolecular forces &amp; types of attraction</li></ul> <p data-bbox="1220 914 1493 938"><b>XIV. Behavior of Gases</b></p> <ul data-bbox="1270 941 1793 1047" style="list-style-type: none"><li>A. Properties of Gases &amp; Gas laws</li><li>B. Direct and indirect – Inverse Relationships</li><li>C. Ideal Gases</li><li>D. Gases – Mixtures &amp; Movement – Mathematical</li></ul> <p data-bbox="1119 1050 1241 1075">Calculations</p> <p data-bbox="1220 1078 1598 1102"><b>XV. Water and Aqueous Systems</b></p> <ul data-bbox="1270 1105 1818 1187" style="list-style-type: none"><li>A. Water and It's Properties</li><li>B. Homogeneous &amp; Heterogeneous Aqueous Systems<ul data-bbox="1325 1154 1562 1187" style="list-style-type: none"><li>1. Hydrate Compounds</li></ul></li></ul> <p data-bbox="1220 1190 1409 1214"><b>XVI. Solutions</b></p> <ul data-bbox="1270 1218 1839 1242" style="list-style-type: none"><li>A. Properties, Concentrations and Colligative Properties</li></ul> <p data-bbox="1220 1245 1509 1269"><b>XVII. Thermochemistry</b></p> <ul data-bbox="1270 1273 1772 1321" style="list-style-type: none"><li>A. Measuring and Expressing Enthalpy Changes</li><li>B. Heat in Changes of State</li></ul> <p data-bbox="1220 1325 1514 1349"><b>XVIII. Acids, Bases, Salts</b></p> <ul data-bbox="1270 1352 1614 1482" style="list-style-type: none"><li>A. Acid – Base Theories</li><li>B. Hydrogen Ions and Acidity</li><li>C. Strength of Acids and Bases</li><li>D. Neutralization Reactions</li><li>E. Salts in Solutions</li></ul> <p data-bbox="1419 1485 1621 1510"><b>BENCHMARK IV</b></p>